

QUIZEN –Structure of the Atoms (9C04)

Learning Level 1	Learning Level 2	Learning Level 3
Q - Remembering (knowledge-based questions) U - Understanding (comprehension-based questions)	I - Applying (application-based questions) Z - Analyzing (analysis-based questions)	E - Evaluating (evaluation-based questions) N - Creating (creation-based questions)

Learning Level 1

1. Define an atom.
2. Who discovered electrons?
3. What is the mass of a proton?
4. Define Atomic Number.
5. What is the charge on a neutron?

Learning Level 2

6. Explain how J.J. Thomson's Cathode Ray Experiment led to the discovery of electrons.
7. Draw and label the Bohr Model of an atom with an atomic number of 12.
8. What is the difference between Atomic Number and Mass Number?
9. What is the significance of Rutherford's Gold Foil Experiment in the study of atomic structure?
10. Calculate the number of neutrons in an atom with an atomic number of 22 and a mass number of 47.

Learning Level 3

11. Evaluate the limitations of Bohr's Atomic Model.
12. Compare and contrast the Plum Pudding Model and the Rutherford Atomic Model.
13. Analyze how the discovery of subatomic particles affected the understanding of atomic structure.
14. What are isotopes? How do they differ from each other?
15. Assess the contributions of Dalton, Thomson, Rutherford, and Bohr in the development of atomic models.